

The University of the State of New York
REGENTS HIGH SCHOOL EXAMINATION

CHEMISTRY

Thursday, June 22, 2000 — 9:15 a.m. to 12:15 p.m., only

The last page of the booklet is the answer sheet. Fold the last page along the perforations and, slowly and carefully, tear off the answer sheet. Then fill in the heading of your answer sheet.

All of your answers are to be recorded on the separate answer sheet. For each question, decide which of the choices given is the best answer. Then on the answer sheet, in the row of numbers for that question, circle with pencil the number of the choice that you have selected. The sample below is an example of the first step in recording your answers.

SAMPLE: (1) 2 3 4

If you wish to change an answer, erase your first penciled circle and then circle with pencil the number of the answer you want. After you have completed the examination and you have decided that all of the circled answers represent your best judgment, signal a proctor and turn in all examination material except your answer sheet. Then and only then, place an X in ink in each penciled circle. Be sure to mark only one answer with an X in ink for each question. No credit will be given for any question with two or more X's marked. The sample below indicates how your final choice should be marked with an X in ink.

SAMPLE: (X) 2 3 4

The "Reference Tables for Chemistry," which you may need to answer some questions in this examination, are supplied separately. Be certain you have a copy of these reference tables before you begin the examination.

When you have completed the examination, you must sign the statement printed at the end of the answer sheet, indicating that you had no unlawful knowledge of the questions or answers prior to the examination and that you have neither given nor received assistance in answering any of the questions during the examination. Your answer sheet cannot be accepted if you fail to sign this declaration.

DO NOT OPEN THIS EXAMINATION BOOKLET UNTIL THE SIGNAL IS GIVEN.

Part I

Answer all 56 questions in this part. [65]

Directions (1–56): For each statement or question, select the word or expression that, of those given, best completes the statement or answers the question. Record your answer on the separate answer sheet in accordance with the directions on the front page of this booklet.

1 The formula Al_2S_3 represents

- (1) an element
- (2) a binary compound
- (3) a ternary compound
- (4) a mixture

2 As the temperature of a gas is increased from 0°C to 10°C at constant pressure, the volume of the gas will

- (1) increase by $\frac{1}{273}$
- (2) increase by $\frac{10}{273}$
- (3) decrease by $\frac{1}{273}$
- (4) decrease by $\frac{10}{273}$

3 Water boils at 90°C when the pressure exerted on the liquid is equal to

- (1) 50.0 torr
- (2) 100.0 torr
- (3) 525.8 torr
- (4) 760.0 torr

4 Which species readily sublimates at room temperature?

- (1) $\text{CO}_2(\text{s})$
- (2) $\text{CO}_2(\ell)$
- (3) $\text{CO}_2(\text{g})$
- (4) $\text{CO}_2(\text{aq})$

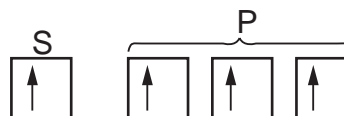
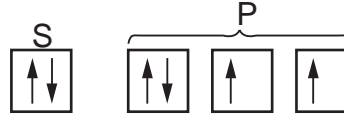
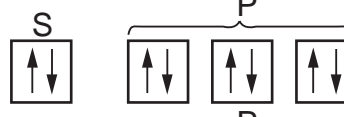

5 Which statement best describes all compounds?

- (1) They can be decomposed by chemical change.
- (2) They can be decomposed by physical means.
- (3) They contain at least three elements.
- (4) They contain ionic bonds.

6 What is the electron configuration of a Mn atom in the ground state?

- (1) $1s^2 2s^2 2p^6 3s^2$
- (2) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^2$
- (3) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1 4p^1$
- (4) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^7$

7 Which orbital notation correctly represents a noble gas in the ground state?

- (1) 
- (2) 
- (3) 
- (4) 

8 Which type of radiation has zero mass and zero charge?

- (1) alpha
- (2) beta
- (3) neutron
- (4) gamma

9 What is the total number of protons and neutrons in an atom of ${}^{86}_{37}\text{Rb}$?

- (1) 37
- (2) 49
- (3) 86
- (4) 123

10 What is the total number of valence electrons in an atom of boron in the ground state?

- (1) 1
- (2) 7
- (3) 3
- (4) 5

11 What causes the emission of radiant energy that produces characteristic spectral lines?

- (1) neutron absorption by the nucleus
- (2) gamma ray emission from the nucleus
- (3) movement of electrons to higher energy levels
- (4) return of electrons to lower energy levels

12 Which atom in the ground state has three half-filled orbitals?

- (1) P (3) Al
(2) Si (4) Li

13 Which particles may be gained, lost, or shared by an atom when it forms a chemical bond?

- (1) protons (3) neutrons
(2) electrons (4) nucleons

14 Which molecular formula is correctly paired with its corresponding empirical formula?

- (1) CO₂ and CO (3) C₆H₆ and C₂H₂
(2) C₂H₂ and CH₂ (4) P₄O₁₀ and P₂O₅

15 Which of the following elements has the strongest attraction for electrons?

- (1) boron (3) oxygen
(2) aluminum (4) sulfur

16 The table below shows four compounds and the boiling point of each.

Compound	Boiling Point
H ₂ O	100.°C
H ₂ S	-60.7°C
H ₂ Se	-41.5°C
H ₂ Te	-2.2°C

Which type of molecular attraction accounts for the high boiling point of H₂O?

- (1) molecule-ion
(2) ion-ion
(3) hydrogen bonding
(4) van der Waals forces

17 Which elements are both classified as metalloids?

- (1) Ge and As (3) B and C
(2) Bi and Po (4) Si and P

18 Which electron dot diagram represents a molecule that has a polar covalent bond?

- (1) $\text{H} \times \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}} \times$ (3) $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}} \times \overset{\times\times}{\underset{\times\times}{\text{Cl}}} \times$
(2) $\text{Li}^+ \left[\times \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}} \times \right]^-$ (4) $\text{K}^+ \left[\times \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}} \times \right]^-$

19 What is the total number of moles of atoms present in 1 gram formula mass of Pb(C₂H₃O₂)₂?

- (1) 9 (3) 3
(2) 14 (4) 15

20 Elements in the Periodic Table are arranged according to their

- (1) atomic number (3) relative activity
(2) atomic mass (4) relative size

21 Which Group 15 element exists as a diatomic molecule at STP?

- (1) phosphorus (3) bismuth
(2) nitrogen (4) arsenic

22 Which element reacts vigorously with water?

- (1) Zn (3) Fe
(2) Cu (4) Li

23 Atoms of metals tend to

- (1) lose electrons and form negative ions
(2) lose electrons and form positive ions
(3) gain electrons and form negative ions
(4) gain electrons and form positive ions

24 Which halogen is a solid at STP?

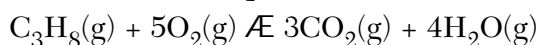
- (1) Br₂ (3) Cl₂
(2) F₂ (4) I₂

25 What is the total number of molecules in a 0.5-mole sample of He gas?

- (1) 6×10^{23} (3) 3×10^{23}
(2) 2×10^{23} (4) 4×10^{23}

- 26 What occurs as the atomic number of the elements in Period 2 increases?
- (1) The nuclear charge of each successive atom decreases, and the covalent radius decreases.
 - (2) The nuclear charge of each successive atom decreases, and the covalent radius increases.
 - (3) The nuclear charge of each successive atom increases, and the covalent radius decreases.
 - (4) The nuclear charge of each successive atom increases, and the covalent radius increases.

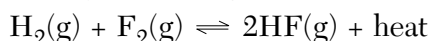
27 Given the balanced equation:



What is the total number of liters of $\text{CO}_2(\text{g})$ produced when 20.0 liters of $\text{O}_2(\text{g})$ are completely consumed?

- (1) 12.0 L
 - (2) 22.4 L
 - (3) 3.00 L
 - (4) 5.00 L
- 28 Given the balanced equation:
- $$\text{Fe}(\text{s}) + \text{CuSO}_4(\text{aq}) \rightarrow \text{FeSO}_4(\text{aq}) + \text{Cu}(\text{s})$$
- What total mass of iron is necessary to produce 1.00 mole of copper?
- (1) 26.0 g
 - (2) 55.8 g
 - (3) 112 g
 - (4) 192 g
- 29 The percent by mass of nitrogen in NH_4NO_3 is closest to
- (1) 15%
 - (2) 20.%
 - (3) 35%
 - (4) 60.%
- 30 What is the molarity of a solution that contains 40. grams of NaOH in 0.50 liter of solution?
- (1) 1.0 M
 - (2) 2.0 M
 - (3) 0.50 M
 - (4) 0.25 M

31 Given the system at equilibrium:

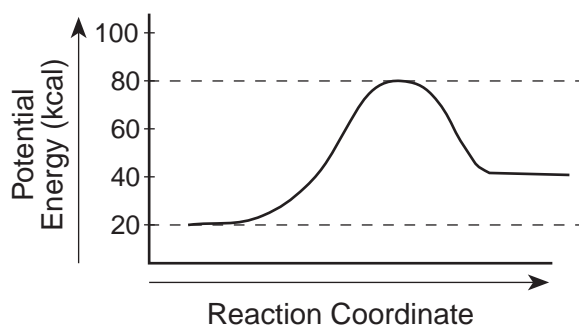


Which change will *not* shift the point of equilibrium?

- (1) changing the pressure
- (2) changing the temperature
- (3) changing the concentration of $\text{H}_2(\text{g})$
- (4) changing the concentration of $\text{HF}(\text{g})$

- 32 An increase in the surface area of reactants in a heterogeneous reaction will result in
- (1) a decrease in the rate of the reaction
 - (2) an increase in the rate of the reaction
 - (3) a decrease in the heat of reaction
 - (4) an increase in the heat of reaction

33 A potential energy diagram of a chemical reaction is shown below.



What is the difference between the potential energy of the reactants and the potential energy of the products?

- (1) 20. kcal
 - (2) 40. kcal
 - (3) 60. kcal
 - (4) 80. kcal
- 34 Given the reaction:
- $$\text{Zn}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq}) + \text{H}_2(\text{g})$$
- If the concentration of the $\text{HCl}(\text{aq})$ is increased, the frequency of reacting collisions will
- (1) decrease, producing a decrease in the reaction rate
 - (2) decrease, producing an increase in the reaction rate
 - (3) increase, producing a decrease in the reaction rate
 - (4) increase, producing an increase in the reaction rate

- 35 Two reactant particles collide with proper orientation. The collision will be effective if the particles have
- (1) high activation energy
 - (2) high ionization energy
 - (3) sufficient kinetic energy
 - (4) sufficient potential energy

36 Based on Reference Table D, what change will cause the solubility of $\text{KNO}_3(\text{s})$ to increase?

- (1) decreasing the pressure
- (2) increasing the pressure
- (3) decreasing the temperature
- (4) increasing the temperature

37 Which of the following ionization constants (K_a) represents the strongest acid?

- (1) $K_a = 1 \times 10^{-14}$
- (2) $K_a = 1 \times 10^{-7}$
- (3) $K_a = 1 \times 10^{-4}$
- (4) $K_a = 1 \times 10^{-2}$

38 Based on Reference Table L, which of the following compounds is the *weakest* electrolyte?

- (1) HI
- (2) HNO_3
- (3) H_3PO_4
- (4) H_2SO_4

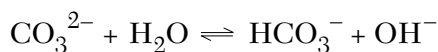
39 Based on Reference Table L, which substance can function only as a Brønsted-Lowry acid?

- (1) HCl
- (2) HSO_4^-
- (3) HCO_3^-
- (4) NH_3

40 If the pH of a solution is 9, the solution is

- (1) acidic, which turns phenolphthalein pink
- (2) acidic, which turns phenolphthalein colorless
- (3) basic, which turns phenolphthalein pink
- (4) basic, which turns phenolphthalein colorless

41 Given the reaction:



Which species is the strongest conjugate base?

- (1) CO_3^{2-}
- (2) H_2O
- (3) HCO_3^-
- (4) OH^-

42 What is the pH of a 0.001 M KOH solution?

- (1) 14
- (2) 11
- (3) 3
- (4) 7

43 Which simple oxidation-reduction reaction is *not* correctly balanced?

- (1) $\text{Sn}(\text{s}) + \text{Cu}^{2+}(\text{aq}) \rightleftharpoons \text{Cu}(\text{s}) + \text{Sn}^{2+}(\text{aq})$
- (2) $\text{Ni}(\text{s}) + \text{Sn}^{2+}(\text{aq}) \rightleftharpoons \text{Sn}(\text{s}) + \text{Ni}^{2+}(\text{aq})$
- (3) $2\text{I}^-(\text{aq}) + \text{Fe}^{3+}(\text{aq}) \rightleftharpoons \text{Fe}^{2+}(\text{aq}) + \text{I}_2(\text{s})$
- (4) $2\text{I}^-(\text{aq}) + \text{Hg}^{2+}(\text{aq}) \rightleftharpoons \text{Hg}(\ell) + \text{I}_2(\text{s})$

44 In the reaction $\text{H}_2\text{S} + \text{NH}_3 \rightleftharpoons \text{NH}_4^+ + \text{HS}^-$, the two Brønsted-Lowry bases are

- (1) NH_3 and HS^-
- (2) NH_3 and NH_4^+
- (3) H_2S and NH_3
- (4) H_2S and HS^-

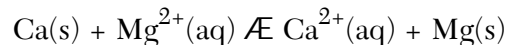
45 A student wishes to set up an electrochemical cell. The following list of materials and equipment will be used:

- two 250-mL beakers
- wire
- one piece of Zn metal
- 125 mL of 0.10 M $\text{Zn}(\text{NO}_3)_2$
- voltmeter
- switch
- one piece of Pb metal
- 125 mL of 0.10 M $\text{Pb}(\text{NO}_3)_2$

For the cell to operate properly, the student will also need

- (1) an anode
- (2) a cathode
- (3) an external path for electrons
- (4) a salt bridge

46 Given the cell reaction:



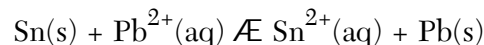
Which substance is oxidized?

- (1) $\text{Ca}(\text{s})$
- (2) $\text{Mg}^{2+}(\text{aq})$
- (3) $\text{Ca}^{2+}(\text{aq})$
- (4) $\text{Mg}(\text{s})$

47 Chlorine has an oxidation state of +3 in the compound

- (1) HClO
- (2) HClO_2
- (3) HClO_3
- (4) HClO_4

48 Given the cell reaction:



The reduction half-reaction for this cell is

- (1) $\text{Pb}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Pb}(\text{s})$
- (2) $\text{Pb}(\text{s}) \rightleftharpoons \text{Pb}^{2+}(\text{aq}) + 2\text{e}^-$
- (3) $\text{Sn}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Sn}(\text{s})$
- (4) $\text{Sn}(\text{s}) \rightleftharpoons \text{Sn}^{2+}(\text{aq}) + 2\text{e}^-$

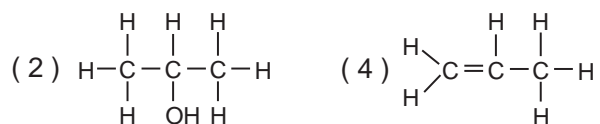
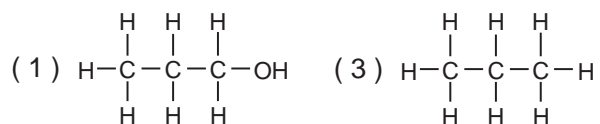
49 Proteins are produced through the process of

- (1) addition
- (2) substitution
- (3) polymerization
- (4) combustion

50 What are the products of a fermentation reaction?

- (1) an alcohol and carbon monoxide
- (2) an alcohol and carbon dioxide
- (3) a salt and water
- (4) a salt and an acid

51 Which structural formula represents a saturated hydrocarbon?



52 Which structural formula represents 1,1-dibromopropane?



53 The principal products of saponification, a reaction between a fat and a base, are soap and

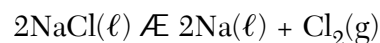
- (1) water
- (2) glycerol
- (3) carbon dioxide
- (4) ethyl alcohol

Note that questions 54 through 56 have only three choices.

54 As a solid substance absorbs heat at its melting point, the melting point will

- (1) decrease
- (2) increase
- (3) remain the same

55 Given the redox reaction:



As the Cl^- is oxidized, the oxidation number of chlorine will

- (1) decrease
- (2) increase
- (3) remain the same

56 As energy is released during the formation of a bond, the stability of the chemical system generally will

- (1) decrease
- (2) increase
- (3) remain the same

Part II

This part consists of twelve groups, each containing five questions. Each group tests a major area of the course. Choose seven of these twelve groups. Be sure that you answer all five questions in each group chosen. Record the answers to these questions on the separate answer sheet in accordance with the directions on the front page of this booklet. [35]

Group 1 — Matter and Energy

If you choose this group, be sure to answer questions 57–61.

- 57 Which term represents a form of energy?
(1) heat (2) degree (3) kilocalorie (4) temperature
- 58 Which change of phase is exothermic?
(1) solid to liquid (2) gas to liquid (3) solid to gas (4) liquid to gas
- 59 A gas sample consisting of 2 moles of hydrogen and 1 mole of oxygen is collected over water at 29°C and 750 torr. What is the partial pressure of the hydrogen in the sample?
(1) 240 torr (2) 480 torr (3) 720 torr (4) 750 torr
- 60 What is the equilibrium temperature of an ice-water mixture at a pressure of 1 atmosphere?
(1) 0°C (2) 32°C (3) 100°C (4) 273°C
- 61 The list below shows four samples: *A*, *B*, *C*, and *D*.
(A) HCl(aq)
(B) NaCl(aq)
(C) HCl(g)
(D) NaCl(s)
- Which samples are substances?
(1) *A* and *B* (2) *A* and *C* (3) *C* and *B* (4) *C* and *D*

Group 2 — Atomic Structure

If you choose this group, be sure to answer questions 62–66.

- 62 Which subatomic particles have a mass of approximately 1 atomic mass unit each?
(1) proton and electron (2) proton and neutron (3) neutron and positron (4) electron and positron
- 63 Which atoms are isotopes of the same element?
(1) ${}_{12}^{24}X$ and ${}_{12}^{25}X$ (2) ${}_{10}^{20}X$ and ${}_{11}^{20}X$ (3) ${}_{15}^{31}X$ and ${}_{16}^{32}X$ (4) ${}_{19}^{31}X$ and ${}_{19}^{31}X$
- 64 What is the total number of grams of a 32-gram sample of ${}^{32}\text{P}$ remaining after 71.5 days of decay?
(1) 1.0 g (2) 2.0 g (3) 8.0 g (4) 4.0 g
- 65 Experiments with gold foil indicated that atoms
(1) usually have a uniform distribution of positive charges
(2) usually have a uniform distribution of negative charges
(3) contain a positively charged, dense center
(4) contain a negatively charged, dense center
- 66 What is the total number of completely filled sublevels found in an atom of krypton in the ground state?
(1) 10 (2) 2 (3) 8 (4) 4

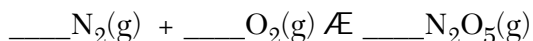
Group 3 — Bonding

If you choose this group, be sure to answer questions 67–71.

67 Which atom will form an ionic bond with a Br atom?

- (1) N (3) O
(2) Li (4) C

68 Given the unbalanced equation:



When the equation is balanced using *smallest* whole numbers, the coefficient of $\text{N}_2(\text{g})$ will be

- (1) 1 (3) 5
(2) 2 (4) 4

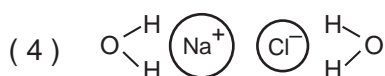
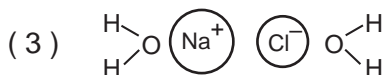
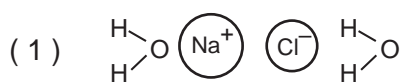
69 Which molecule is polar and contains polar bonds?

- (1) CCl_4 (3) N_2
(2) CO_2 (4) NH_3

70 The *strongest* van der Waals forces of attraction exist between molecules of

- (1) I_2 (3) Cl_2
(2) Br_2 (4) F_2

71 Which diagram best illustrates the ion-molecule attractions that occur when the ions of $\text{NaCl}(\text{s})$ are added to water?



Group 4 — Periodic Table

If you choose this group, be sure to answer questions 72–76.

72 An element has a first ionization energy of 314 kilocalories/mole and an electronegativity of 3.5. It is classified as a

- (1) metal (3) metalloid
(2) nonmetal (4) halogen

73 At which location in the Periodic Table would the most active metallic element be found?

- (1) in Group 1 at the top
(2) in Group 1 at the bottom
(3) in Group 17 at the top
(4) in Group 17 at the bottom

74 Which set of properties is most characteristic of transition elements?

- (1) colorless ions in solution, multiple positive oxidation states
(2) colorless ions in solution, multiple negative oxidation states
(3) colored ions in solution, multiple positive oxidation states
(4) colored ions in solution, multiple negative oxidation states

75 An atom with the electron configuration $1s^2 2s^2 2p^6 3s^2$ would most likely

- (1) decrease in size as it forms a positive ion
(2) increase in size as it forms a positive ion
(3) decrease in size as it forms a negative ion
(4) increase in size as it forms a negative ion

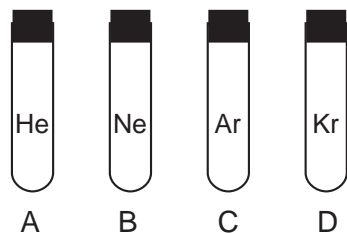
76 The properties of carbon are expected to be most similar to those of

- (1) boron (3) silicon
(2) aluminum (4) phosphorus

Group 5 — Mathematics of Chemistry

If you choose this group, be sure to answer questions 77–81.

- 77 The stoppered tubes below, labeled A through D, each contain a different gas.



When the tubes are unstoppered at the same time and under the same conditions of temperature and pressure, from which tube will gas diffuse at the fastest rate?

- (1) A (3) C
(2) B (4) D
- 78 A compound whose empirical formula is NO_2 could have a molecular mass of
- (1) 23 (3) 92
(2) 39 (4) 120

- 79 The density of a gas is 1.43 grams per liter at STP. The mass of 1 mole of this gas is equal to

- (1) 1.43 g (3) 22.4 g
(2) 15.7 g (4) 32.0 g

- 80 What is the total number of kilocalories required to boil 100. grams of water at 100°C and 1 atmosphere? [Refer to Reference Table A.]

- (1) 1.80 kcal (3) 53.9 kcal
(2) 18.0 kcal (4) 539 kcal

- 81 Which property of a distilled water solution will *not* be affected by adding 50 mL of $\text{CH}_3\text{OH}(\ell)$ to 100 mL of the water solution at 25°C ?

- (1) conductivity (3) freezing point
(2) vapor pressure (4) boiling point

Group 6 — Kinetics and Equilibrium

If you choose this group, be sure to answer questions 82–86.

82 Based on Reference Table *M*, which compound has a K_{sp} closest to the K_{sp} of PbCrO_4 ?

- (1) Ag_2CrO_4 (3) ZnCO_3
(2) AgBr (4) PbCl_2

83 Adding a catalyst to a chemical reaction changes the rate of reaction by causing

- (1) a decrease in the activation energy
(2) an increase in the activation energy
(3) a decrease in the heat of reaction
(4) an increase in the heat of reaction

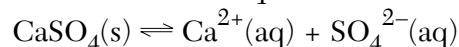
84 Which change in a sample of water is accompanied by the greatest increase in entropy?

- (1) $\text{H}_2\text{O}(\ell)$ at 100°C is changed to $\text{H}_2\text{O}(\text{g})$ at 200°C .
(2) $\text{H}_2\text{O}(\text{g})$ at 100°C is changed to $\text{H}_2\text{O}(\text{g})$ at 200°C .
(3) $\text{H}_2\text{O}(\text{s})$ at -100°C is changed to $\text{H}_2\text{O}(\text{s})$ at 0°C .
(4) $\text{H}_2\text{O}(\text{s})$ at -100°C is changed to $\text{H}_2\text{O}(\ell)$ at 0°C .

85 According to Reference Table *G*, which reaction spontaneously forms a compound from its elements?

- (1) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
(2) $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g})$
(3) $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$
(4) $\text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$

86 Given the solution at equilibrium:



When Na_2SO_4 is added to the system, how will the equilibrium shift?

- (1) The amount of $\text{CaSO}_4(\text{s})$ will decrease, and the concentration of $\text{Ca}^{2+}(\text{aq})$ will decrease.
(2) The amount of $\text{CaSO}_4(\text{s})$ will decrease, and the concentration of $\text{Ca}^{2+}(\text{aq})$ will increase.
(3) The amount of $\text{CaSO}_4(\text{s})$ will increase, and the concentration of $\text{Ca}^{2+}(\text{aq})$ will decrease.
(4) The amount of $\text{CaSO}_4(\text{s})$ will increase, and the concentration of $\text{Ca}^{2+}(\text{aq})$ will increase.

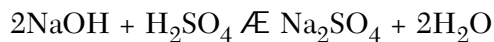
Group 7 — Acids and Bases

If you choose this group, be sure to answer questions 87–91.

87 When HCl is dissolved in water, the only positive ion present in the solution is the

- (1) hydrogen ion (3) hydride ion
(2) hydroxide ion (4) chloride ion

88 Given the reaction:



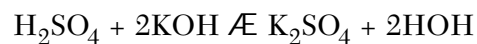
How many milliliters of 1 M NaOH are needed to exactly neutralize 100 milliliters of 1 M H_2SO_4 ?

- (1) 50 mL (3) 300 mL
(2) 200 mL (4) 400 mL

89 According to the Brønsted-Lowry theory, H_2O is considered to be a base when it

- (1) donates an electron (3) donates a proton
(2) accepts an electron (4) accepts a proton

90 Given the neutralization reaction:



Which compound is a salt?

- (1) KOH (3) K_2SO_4
(2) H_2SO_4 (4) HOH

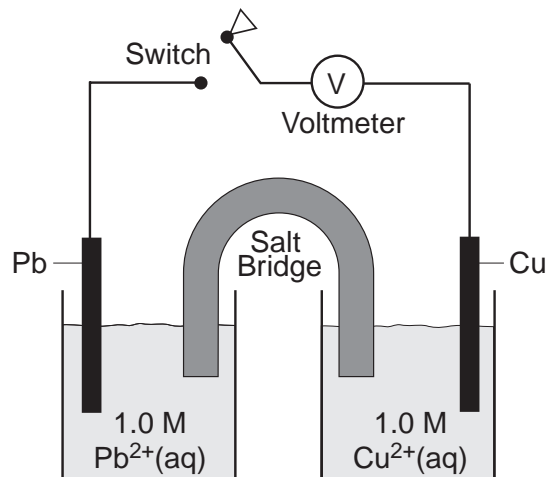
91 Which acid-base pair will always undergo a reaction that produces a neutral solution?

- (1) a weak acid and a weak base
(2) a weak acid and a strong base
(3) a strong acid and a weak base
(4) a strong acid and a strong base

Group 8 — Redox and Electrochemistry

If you choose this group, be sure to answer questions 92–96.

Base your answers to questions 92 and 93 on the diagram of a chemical cell and the equation shown below. The reaction occurs at 1 atmosphere and 298 K.



92 When the switch is closed, the cell voltage (E^0) is

- (1) -0.21 V (3) -0.47 V
(2) $+0.21\text{ V}$ (4) $+0.47\text{ V}$

93 Which change occurs when the switch is closed?

- (1) Pb is oxidized, and electrons flow to the Cu electrode.
(2) Pb is reduced, and electrons flow to the Cu electrode.
(3) Cu is oxidized, and electrons flow to the Pb electrode.
(4) Cu is reduced, and electrons flow to the Pb electrode.

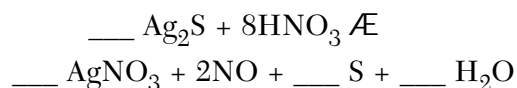
94 Based on Reference Table N, the standard electrode potential for the reduction of gold (III) ions is

- (1) $+1.50\text{ V}$ (3) -0.80 V
(2) $+0.80\text{ V}$ (4) -1.50 V

95 In an electrolytic cell, a negative ion will migrate to and undergo oxidation at the

- (1) anode, which is negatively charged
(2) anode, which is positively charged
(3) cathode, which is negatively charged
(4) cathode, which is positively charged

96 Given the unbalanced equation:



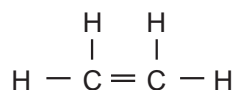
What is the coefficient of Ag_2S when the equation is completely balanced using the *smallest* whole numbers?

- (1) 6 (3) 3
(2) 2 (4) 4

Group 9 — Organic Chemistry

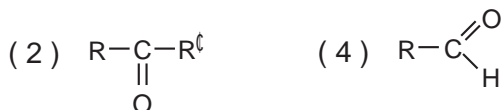
If you choose this group, be sure to answer questions 97–101.

97 Given the compound:



The symbol = represents

- (1) one pair of shared electrons
 - (2) two pairs of shared electrons
 - (3) a single covalent bond
 - (4) a coordinate covalent bond
- 98 An organic compound containing one or more OH groups as the only functional group is classified as an
- (1) aldehyde
 - (2) alcohol
 - (3) ester
 - (4) ether
- 99 The reaction during which monomers are combined and water is released is called
- (1) saponification
 - (2) neutralization
 - (3) addition polymerization
 - (4) condensation polymerization
- 100 One molecule of glycerol contains a total of
- (1) two –OH groups
 - (2) two –CH₃ groups
 - (3) three –OH groups
 - (4) three –CH₃ groups
- 101 What is the general formula for an ether?



Group 10 — Applications of Chemical Principles

If you choose this group, be sure to answer questions 102–106.

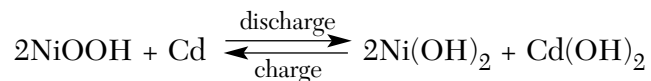
102 Which type of chemical reaction occurs in a lead-acid battery?

- (1) addition
- (2) substitution
- (3) esterification
- (4) oxidation-reduction

103 Which metals are obtained by electrolysis of their fused salts?

- (1) K and Ca
- (2) K and Cr
- (3) Cu and Zn
- (4) Cu and Hg

104 Given the redox reaction:



Which species is oxidized during discharge?

- (1) Cd
- (2) Cd²⁺
- (3) Ni(OH)₃
- (4) Ni(OH)₂

105 Petroleum is primarily a mixture of

- (1) alcohol molecules
- (2) ester molecules
- (3) hydrocarbon molecules
- (4) organic acid molecules

106 By which process is petroleum separated into its components according to their different boiling points?

- (1) contact process
 - (2) Haber process
 - (3) fractional distillation
 - (4) cracking
-

Group 11 — Nuclear Chemistry

If you choose this group, be sure to answer questions 107–111.

107 Organic molecules react to form a product. These reactions may be studied by using

- (1) Sr-90
- (2) Co-60
- (3) N-16
- (4) C-14

108 In a fission reactor, the speed of the neutrons may be decreased by

- (1) a moderator
- (2) an accelerator
- (3) a fuel rod
- (4) shielding

109 Which statement explains why fusion reactions are difficult to initiate?

- (1) Positive nuclei attract each other.
- (2) Positive nuclei repel each other.
- (3) Neutrons prevent nuclei from getting close enough to fuse.
- (4) Electrons prevent nuclei from getting close enough to fuse.

110 A particle accelerator is used to provide charged particles with sufficient

- (1) kinetic energy to penetrate a nucleus
- (2) kinetic energy to penetrate an electron cloud
- (3) potential energy to penetrate a nucleus
- (4) potential energy to penetrate an electron cloud

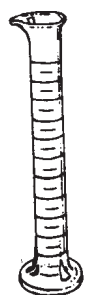
111 In which reaction is mass converted to energy by the process of fission?

- (1) ${}^{14}_7\text{N} + {}^1_0\text{n} \rightarrow {}^{14}_6\text{C} + {}^1_1\text{H}$
- (2) ${}^{235}_{92}\text{U} + {}^1_0\text{n} \rightarrow {}^{87}_{35}\text{Br} + {}^{146}_{57}\text{La} + 3{}^1_0\text{n}$
- (3) ${}^{226}_{88}\text{Ra} \rightarrow {}^{222}_{86}\text{Rn} + {}^4_2\text{He}$
- (4) ${}^2_1\text{H} + {}^2_1\text{H} \rightarrow {}^4_2\text{He}$

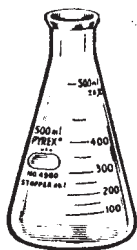
Group 12 — Laboratory Activities

If you choose this group, be sure to answer questions 112–116.

112 Which piece of glassware is used for accurately measuring volumes of an acid and a base during a titration?



(1)



(2)



(3)



(4)

113 The results of testing a colorless solution with three indicators are shown in the table below.

Indicator	Result
red litmus	blue
blue litmus	blue
phenolphthalein	pink

Which formula could represent the solution tested?

- (1) NaOH(aq) (3) C₆H₁₂O₆(aq)
 (2) HCl(aq) (4) C₁₂H₂₂O₁₁(aq)

114 What is the product of (2.324 cm × 1.11 cm) expressed to the correct number of significant figures?

- (1) 2.58 cm² (3) 2.5796 cm²
 (2) 2.5780 cm² (4) 2.57964 cm²

115 A student determined the percentage of water of hydration in BaCl₂•2H₂O by using the data in the table below.

Quantity Measured	Value Obtained
mass of BaCl ₂ •2H ₂ O	3.80 grams
mass of BaCl ₂	3.20 grams
% of water calculated	15.79%

The accepted percentage value for the water of hydration is 14.75%. What is the student's percent error?

- (1) 1.04% (3) 6.59%
 (2) 6.00% (4) 7.05%

116 By which process is a precipitate most easily separated from the liquid in which it is suspended?

- (1) neutralization (3) condensation
 (2) distillation (4) filtration

Part II (35 credits)

Answer the questions in only seven of the twelve groups in this part. Be sure to mark the answers to the groups of questions you choose in accordance with the instructions on the front cover of the test booklet. Leave blank the five groups of questions you do not choose to answer.

Tear Here

Group 1 Matter and Energy					
57	1	2	3	4	
58	1	2	3	4	
59	1	2	3	4	
60	1	2	3	4	
61	1	2	3	4	

Group 2 Atomic Structure					
62	1	2	3	4	
63	1	2	3	4	
64	1	2	3	4	
65	1	2	3	4	
66	1	2	3	4	

Group 3 Bonding					
67	1	2	3	4	
68	1	2	3	4	
69	1	2	3	4	
70	1	2	3	4	
71	1	2	3	4	

Group 4 Periodic Table					
72	1	2	3	4	
73	1	2	3	4	
74	1	2	3	4	
75	1	2	3	4	
76	1	2	3	4	

Group 5 Mathematics of Chemistry					
77	1	2	3	4	
78	1	2	3	4	
79	1	2	3	4	
80	1	2	3	4	
81	1	2	3	4	

Group 6 Kinetics and Equilibrium					
82	1	2	3	4	
83	1	2	3	4	
84	1	2	3	4	
85	1	2	3	4	
86	1	2	3	4	

Group 7 Acids and Bases					
87	1	2	3	4	
88	1	2	3	4	
89	1	2	3	4	
90	1	2	3	4	
91	1	2	3	4	

Group 8 Redox and Electrochemistry					
92	1	2	3	4	
93	1	2	3	4	
94	1	2	3	4	
95	1	2	3	4	
96	1	2	3	4	

Group 9 Organic Chemistry					
97	1	2	3	4	
98	1	2	3	4	
99	1	2	3	4	
100	1	2	3	4	
101	1	2	3	4	

Group 10 Applications of Chemical Principles					
102	1	2	3	4	
103	1	2	3	4	
104	1	2	3	4	
105	1	2	3	4	
106	1	2	3	4	

Group 11 Nuclear Chemistry					
107	1	2	3	4	
108	1	2	3	4	
109	1	2	3	4	
110	1	2	3	4	
111	1	2	3	4	

Group 12 Laboratory Activities					
112	1	2	3	4	
113	1	2	3	4	
114	1	2	3	4	
115	1	2	3	4	
116	1	2	3	4	

I do hereby affirm, at the close of this examination, that I had no unlawful knowledge of the questions or answers prior to the examination and that I have neither given nor received assistance in answering any of the questions during the examination.

Signature

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The University of the State of New York

REGENTS HIGH SCHOOL EXAMINATION

CHEMISTRY

Thursday, June 22, 2000 — 9:15 a.m. to 12:15 p.m., only

ANSWER SHEET

Student Sex: Male
 Female
 Teacher
 School

Record all of your answers on this answer sheet in accordance with the instructions on the front cover of the test booklet.

Part I (65 credits)

- | | | | | | | | | | | | | | | |
|----|---|---|---|---|----|---|---|---|---|----|---|---|---|---|
| 1 | 1 | 2 | 3 | 4 | 21 | 1 | 2 | 3 | 4 | 41 | 1 | 2 | 3 | 4 |
| 2 | 1 | 2 | 3 | 4 | 22 | 1 | 2 | 3 | 4 | 42 | 1 | 2 | 3 | 4 |
| 3 | 1 | 2 | 3 | 4 | 23 | 1 | 2 | 3 | 4 | 43 | 1 | 2 | 3 | 4 |
| 4 | 1 | 2 | 3 | 4 | 24 | 1 | 2 | 3 | 4 | 44 | 1 | 2 | 3 | 4 |
| 5 | 1 | 2 | 3 | 4 | 25 | 1 | 2 | 3 | 4 | 45 | 1 | 2 | 3 | 4 |
| 6 | 1 | 2 | 3 | 4 | 26 | 1 | 2 | 3 | 4 | 46 | 1 | 2 | 3 | 4 |
| 7 | 1 | 2 | 3 | 4 | 27 | 1 | 2 | 3 | 4 | 47 | 1 | 2 | 3 | 4 |
| 8 | 1 | 2 | 3 | 4 | 28 | 1 | 2 | 3 | 4 | 48 | 1 | 2 | 3 | 4 |
| 9 | 1 | 2 | 3 | 4 | 29 | 1 | 2 | 3 | 4 | 49 | 1 | 2 | 3 | 4 |
| 10 | 1 | 2 | 3 | 4 | 30 | 1 | 2 | 3 | 4 | 50 | 1 | 2 | 3 | 4 |
| 11 | 1 | 2 | 3 | 4 | 31 | 1 | 2 | 3 | 4 | 51 | 1 | 2 | 3 | 4 |
| 12 | 1 | 2 | 3 | 4 | 32 | 1 | 2 | 3 | 4 | 52 | 1 | 2 | 3 | 4 |
| 13 | 1 | 2 | 3 | 4 | 33 | 1 | 2 | 3 | 4 | 53 | 1 | 2 | 3 | 4 |
| 14 | 1 | 2 | 3 | 4 | 34 | 1 | 2 | 3 | 4 | 54 | 1 | 2 | 3 | |
| 15 | 1 | 2 | 3 | 4 | 35 | 1 | 2 | 3 | 4 | 55 | 1 | 2 | 3 | |
| 16 | 1 | 2 | 3 | 4 | 36 | 1 | 2 | 3 | 4 | 56 | 1 | 2 | 3 | |
| 17 | 1 | 2 | 3 | 4 | 37 | 1 | 2 | 3 | 4 | | | | | |
| 18 | 1 | 2 | 3 | 4 | 38 | 1 | 2 | 3 | 4 | | | | | |
| 19 | 1 | 2 | 3 | 4 | 39 | 1 | 2 | 3 | 4 | | | | | |
| 20 | 1 | 2 | 3 | 4 | 40 | 1 | 2 | 3 | 4 | | | | | |

Your answers for Part II should be placed in the proper spaces on the back of this sheet.

FOR TEACHER USE ONLY

Credits

Part I
 (Use table below)

Part II

Total

Rater's Initials:

Part I Credits

Directions to Teacher:

In the table below, draw a circle around the number of right answers and the adjacent number of credits. Then write the number of credits (not the number right) in the space provided above.

No. Right	Credits	No. Right	Credits
56	65	28	41
55	64	27	40
54	63	26	39
53	62	25	39
52	62	24	38
51	61	23	37
50	60	22	36
49	59	21	35
48	58	20	34
47	57	19	33
46	56	18	33
45	56	17	32
44	55	16	31
43	54	15	30
42	53	14	29
41	52	13	27
40	51	12	25
39	51	11	23
38	50	10	21
37	49	9	19
36	48	8	17
35	47	7	14
34	46	6	12
33	45	5	10
32	45	4	8
31	44	3	6
30	43	2	4
29	42	1	2
		0	0

No. right

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